

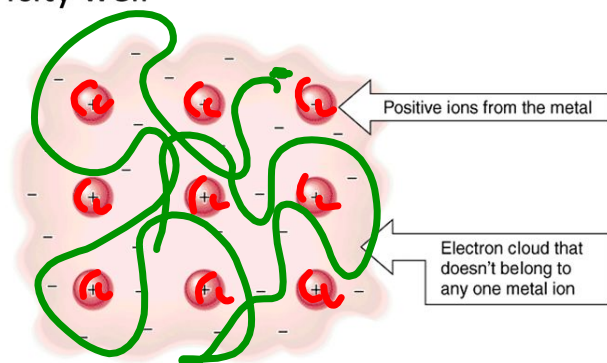
Bond Types *M = metal N = nonmetal*

	Ionic	<u>Covalent</u>	Metallic 
Formula	$M + NM$	$NM + NM$	$M + M$
Valence Electrons	<i>transferred (ions formed)</i>	<i>shared</i>	<i>"sea of free floating electrons"</i>
Melting point	<i>high</i>	<i>low</i>	<i>crazy high</i>
Conducts electricity	<i>only when dissolved.</i>	<i>never</i>	<i>always.</i>



Metallic Bonding

- **"sea of electrons"**: electrons can "float" freely between atoms; allows metals to conduct electricity well

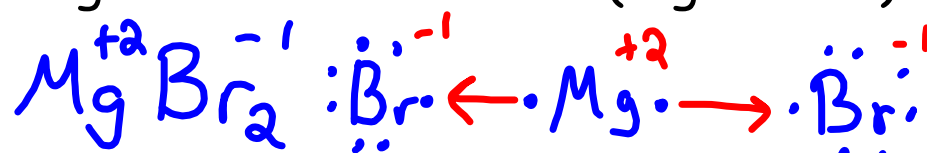


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Drawing Bonds with Lewis Structures

Ionic - Electrons Transferred, Ions formed

1) magnesium and bromine (Mg and Br)



Covalent - Electrons Shared

1) nitrogen triiodide (NI_3)

