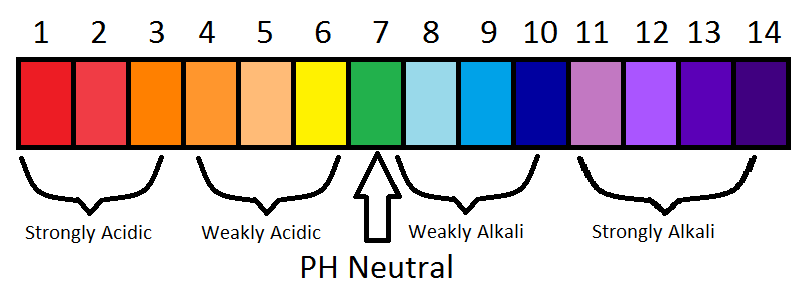
Acid and Base Note Guide

The \_\_\_\_\_\_\_\_ scale measures how acidic or basic something is.



An acid has more \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ions.

A base has more \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ions.

Something neutral, like water, has a pH of \_\_\_\_\_\_\_ . That means it has equal numbers of H+ ions and OH- ions!

(H+) + (OH-) = \_\_\_\_\_\_\_\_\_

The pH scale is logarithmic. That means it increases by factors of \_\_\_\_\_\_\_\_ .

Tomato juice (pH 4) is \_\_\_\_\_\_\_\_\_\_ times more acidic than water (pH 7).

Bleach (pH 12) is \_\_\_\_\_\_\_\_\_\_\_ times more basic than water (pH 7).

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ are weak acids or bases that help keep pH within a narrow range.

Buffers help to keep our blood at a pH of \_\_\_\_\_\_\_\_\_\_ .