JOURNAL 11/17

During our lab how did you know when one of the elements reacted with the hydrochloric acid or the copper chloride?

bubbles, heat, new stuff, change color, disintegrating

4th Hour TAKE OUT YOUR READING GUIDE FOR GRADING

How is a compound different from a element? 11/17

2 elements - ch. combined
element just one thing
" can be found on p. table

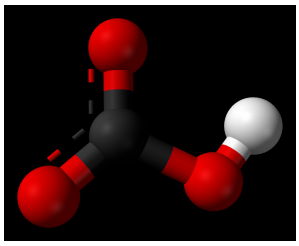
How are they similar?

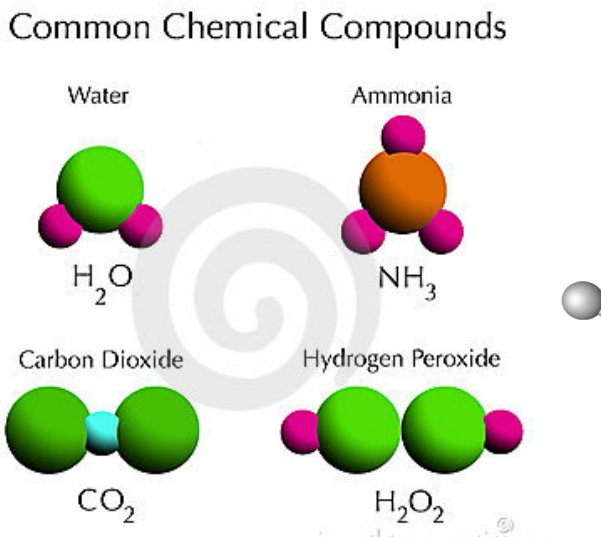
can't be broken physically
both have ph/ch properties
both consist of elements
" " of atoms

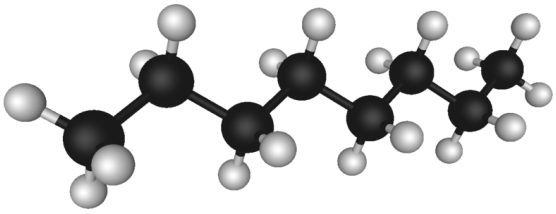
Agenda

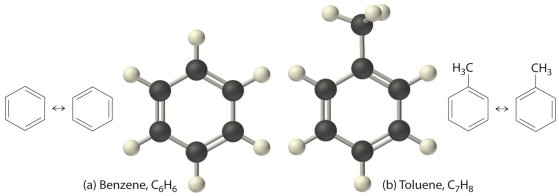
- Discuss Lab
- Reading guide

Compounds: a pure substance made up of two or more atoms that are chemically combined.

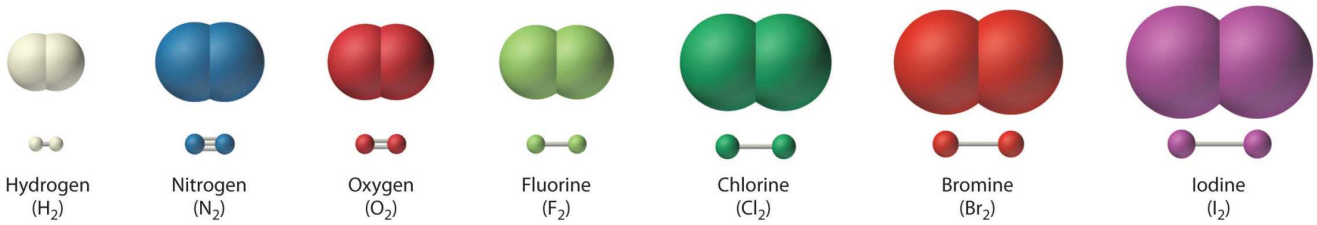




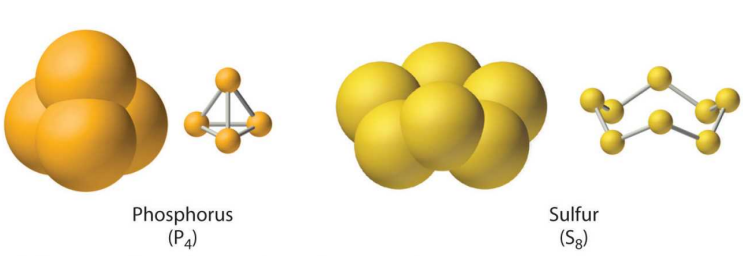




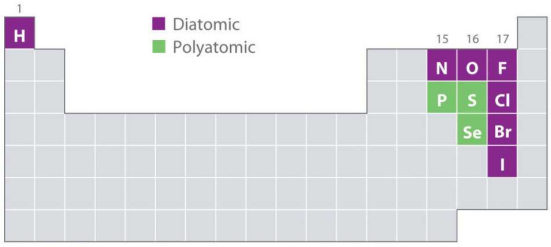
Element or Compound?



(a) Elements that exist as diatomic molecules



(b) Elements that exist as polyatomic molecules



Compound Demonstrations

https://youtu.be/d2geiGKFveE

Na^+ Cl^-

NaCl

1 : 1

Cl^-

H_2O

$H_2O + Cl_2 \rightarrow HCl + O_2$

Discussion

What was up with the Sulfur?

never insoluble

hydrophobic

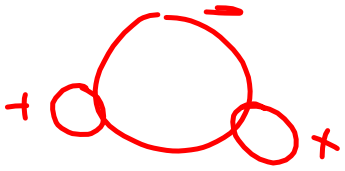
water fear

What reacted with the Copper Chloride?

How do we know?

Al

Zn



S

What about the copper in the acid?

What did react with acid?

$Al + CuCl_2 \rightarrow AlCl_3 + Cu$

$Zn + CuCl_2 \rightarrow ZnCl_2 + Cu$

$Zn + HCl \rightarrow H_2 + ZnCl_2$

Cl — Cl

F- CaF MgF H_2CO_3

$Zn + HCl \rightarrow H_2 + ZnCl_2$

Attachments

- Atomic Theory.ppt
- Atoms, Elements and Properties.pptx
- Properties.pptx
- Element Observation.docx
- Frayer model 2.doc
- Notes on Matter.pptx
- abbreviation in metric system.notebook
- Measurement in the Lab.pdf
- Observation and Inference Only Notes.notebook
- matter notes and journal.notebook
- how to build a data table lesson.notebook
- particles of matter journal and notes.notebook
- States of Matter journal and notes.notebook
- Gas Laws.notebook

Properties.pptx

- What is an Element.pptx

M NM MD Graphic Organizer Class Notes.docx